

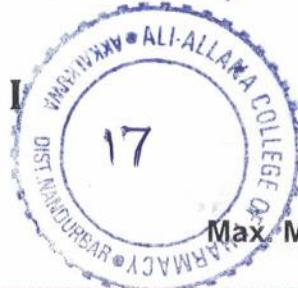
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Max Marks : 85

BP 102 - T

Pharmaceutical Analysis - I (711102)

P. Pages : 3**Time : Three Hours****Instructions to Candidates :**

1. Do not write anything on question paper except Seat No.
2. Graph or diagram should be drawn with the black ink pen being used for writing paper or black HB pencil.
3. Students should note, no supplement will be provided.
4. All questions are compulsory.
5. Draw well labelled diagram wherever necessary.

1. Multiple choice questions.

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- i) Bromatometry involve below oxidizing agent.
 - a) Potassium bromate
 - b) Potassium iodide
 - c) Potassium dichromate
 - d) Iodine
- ii) Which indicator used in complexometric titrations
 - a) Crystal Violet
 - b) Brilliant Green
 - c) Methyl Orange
 - d) None of above
- iii) Which method use in expressing concentration
 - a) Equivalent weight
 - b) Molecular weight
 - c) Ionic mobility
 - d) Normality and Molarity
- iv) Which is type of error in following.
 - a) Accuracy
 - b) Precision
 - c) Significant figure
 - d) Determinate and Indeterminate
- v) Diazotization is explained as.
 - a) Reaction between primary aromatic amine and nitrous acid
 - b) Reaction between cyano and nitric acid.
 - c) Reaction between 920 compounds and HCl
 - d) Reaction between Dibasic amine and acid.
- vi) For gravimetry following filter paper used in analysis.
 - a) Ash value
 - b) Ash less
 - c) Normal paper
 - d) Filter paper
- vii) Iodometry involve
 - a) Titration with potassium iodate (KIO_3)
 - b) Titration with Iodine
 - c) Titration with sodium thiosulphate
 - d) Titration with ceric ammonium sulphate.

- viii) These elements are determined by limit test.
- HCl
 - NaDH
 - Arsenic and Lead
 - Anthranilic acid
- ix) Which titration fall under the category of Argentometric titration.
- Estimation of sodium chloride
 - Estimation of calcium gluconate
 - Iodometry
 - Iodimetry.
- x) Which titration use in complexometric titration.
- Estimation of sodium chloride
 - Estimation of sodium benzoate
 - Estimation of calcium gluconate
 - Estimation of barium sulphate
- xi) Which titration fall under the category of acid base titrations.
- Estimation of sodium chloride
 - Iodometric estimation
 - Estimation of strong, weak acids.
 - Estimation of Magnesium sulphate.
- xii) Oxidation means.
- Loss of electron
 - Gain of electron
 - Loss of proton
 - None of above
- xiii) Iodometry involve titration with
- NaOH
 - HCl
 - Sodium thio sulphate
 - Starch.
- xiv) In Iodometry following substance use as indicator
- Sodium thiosulphate
 - Methylene blue
 - Potassium iodide
 - Starch.
- xv) What do you mean by conductometry?
- Determine potential
 - Determine conductance
 - Determine concentration
 - Determine pH.
- xvi) Which Argentometric titration involve formation of red coloured intermediate ferrocyanide complex.
- Mohr's
 - Fajan's
 - Volhard's
 - Guy Crsal's
- xvii) DME (dropping mercury electrode) involve following reaction.
- Amalgam Formation
 - Diazotization
 - Precipitation
 - None of above
- xviii) Which chelating agent is polydentate
- Ethylene diamine
 - EDTA
 - Both a & b
 - None of above
- xix) Which indicator used in complexometric titration.
- Methylene Blue
 - Crystal Violet
 - Eriochrome black T
 - Methyl Orange
- Which cell use in conductometry.
- Combine glass electrode
 - Calomel electrode
 - Silver chloride electrode
 - Normal hydrogen electrode.



2. Solve any two.

- Discuss the principle involved in conductometry and explain various types of conductometric titrations.
- Discuss the principle involved in gravimetric analysis, Enlist steps involved in it, and explain the concept of co-precipitation with example.
- What do you mean by error. Explain the sources and types of errors in detail.

3. Solve any seven

- Explain iodometry and iodimetry.
- What are primary standard substances. Explain with example
- Explain Mohr's method in detail.
- Add note on masking and demasking agents.
- Describe diazotization titrations.
- Write note on sources of impurities in medicinal agents.
- Discuss various solvents use in Non-aqueous titration.
- Explain theories of indicator add note on universal indicator
- Explain the factors affecting purify of precipitate.

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