

Seat Number

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W-18

08 - 008



BP 104 T  
**Pharmaceutical Inorganic Chemistry**  
(711104)

P. Pages : 3

Time : Three Hours

Max. Marks : 50

Instructions to Candidates :

1. Do not write anything on question paper except Seat No.
2. Graph or diagram should be drawn with the black ink pen being used for writing paper or black HB pencil.
3. Students should note, no supplement will be provided.
4. Figures to the right indicate full marks.

1. Choose the appropriate alternative from those given below and rewrite the correct sentence (all the questions are compulsory). 20

i) .....is responsible for reduction of ferric ions to ferrous ions in the limit test for Iron I.P.

- |                      |                      |
|----------------------|----------------------|
| a) Sulphuric acid    | b) Nitric acid       |
| c) Thioglycolic acid | d) Hydrochloric acid |

ii) ..... is used to prevent precipitation of silver as its carbonate and phosphate in the limit test for chloride.

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|----------------------|-------------------|
| a) Nitric acid       | b) Silver nitrate |
| c) Hydrochloric acid | d) None of these  |

iii) ..... solution of NaCl is isotonic.

- |             |             |
|-------------|-------------|
| a) 0.9% W/v | b) 0.5% W/v |
| c) 1.5% W/v | d) 2.5% W/v |

iv) .....cation is major intracellular electrolyte.

- |                     |                  |
|---------------------|------------------|
| a) $\text{Na}^+$    | b) $\text{K}^+$  |
| c) $\text{Mg}^{++}$ | d) None of these |

v)  $\Delta\text{TF}$  of blood is.....

- |          |          |
|----------|----------|
| a) -0.58 | b) -0.48 |
| c) -0.2  | d) -0.68 |

vi) Which of the following statement is correct for adjustment of tonicity using white Incient Method?

- Solute is used for adjustment of tonicity.
- Solvent is used for adjustment of tonicity.
- Solute and solvent are used for adjustment of tonicity
- None of the above

- vii) Acidic buffer solution contains -----  
 a) Strong acid and its salt.      b) Weak base and its acid.  
 c) Weak acid and its salt.      d) Strong base and its acid.
- viii) Which statement is correct to prove effect of Impurities?  
 a) They leads to toxic effects.  
 b) They decrease therapeutic effect of drugs.  
 c) They may cause problems in formulations.  
 d) All the above.
- ix) the properties of solution which depend on number of particles present in solution is called -----  
 a) Colligative properties      b) Heterogenous properties  
 c) Homogenous properties      d) Osmotic properties.
- x) Addison's disease is responsible for -----  
 a) Hyponatremia      b) Hypernatraemia  
 c) Hypokalemia      d) Hyperkalemia
- xi) ----- Particles posses maximum penetrating power.  
 a) Alpha      b) Beta  
 c) Gamma      d) None of these.
- xii) Agents which are used to prevent microbial spoilage of the preparation is called -----  
 a) Carminatives      b) Preservatives  
 c) Sedatives      d) All of these
- xiii) Iodine is prepared from -----  
 a) Sea-Weeds      b) Child saltpeter  
 c) Brines      d) All of these.
- xiv) Antacids results in decrease bioavailability of -----  
 a) Certain vitamins & Iron      b) Benzodiazepines  
 c) Isoniazid      d) All of these,
- xv) Delayed evacuation of the bowels / Feces is called -----  
 a) Acidity      b) Cathartic  
 c) Constipation      d) Peristalsis.
- xvi) Which of the following statement maps the properties of buffer solutions.  
 a) pH remains constant  
 b) pH does not change  
 c) pH does not alter over specific period  
 d) All of these
- xvii) To prevent dental caries it becomes necessary to use -----  
 a) Sodium Fluoride      b) Gold particles  
 c) Silver nitrate      d) Copper sulphate

xviii) Potassium iodide can be used as/in -----

- a) Expectorant
- b) Thyroid deficiency
- c) Saline diuretic
- d) All of these

xix) Which of the following statement satisfies mechanism of action of Antidotes?

- i) Act by neutralizing Poison
- ii) Antagonistic action
- iii) Converts toxic to less/non-toxic forms
- iv) None of these
- a) Statement i & ii
- b) Statement i, ii & iii
- c) Statement ii & iii
- d) Statement iv

xx) The pH of normal blood ranges between -----

- a) 7.35 and 7.45
- b) 7.15 and 7.25
- c) 7.05 and 7.15
- d) 7.65 and 7.75

2.

Solve **any two** of the following.

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- a) Enlist and explain the various sources and types of Impurities
- b) Explain the mechanism of action of anti microbial agent with suitable example and add a note on Hydrogen peroxide solution.
- c) What do you mean by radioactive substances? Discuss in detail about measurement of radioactivity.

3.

Write short notes on the following **any seven**.

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- a) Principle behind limit test for Arsenic I.
- b) Buffer capacity
- c) Properties of an ideal antacid
- d) Tonicity
- e) Pharmacopoeia
- f) Ammonium chloride
- g) Properties and uses of Copper sulphate.
- h) Dentifrices and role of fluoride
- i) Major Intracellular and Extracellular electrolytes

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